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COMMUNICATION

No more compromise: a facile route towards functionalized surfaces with stable monolayers

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The results reported in this study clearly show that it is possible to easily control the formation of a functional monolayer by spontaneous reduction of an aryldiazonium salt on gold in a single step, mimicking the SAM technique.

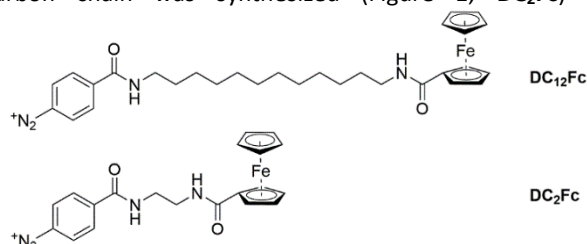
The immobilization of functional organic molecules is a commonly adopted strategy to tailor surface chemistry and produce new specific properties.^{1–3} The choice of the used method depends both on the scientific background of the authors and on the goal (or application) sought. In any case, it is necessary to compromise either on the stability of the layers obtained, the control of their growth, their ease of implementation or sometimes even on the choice of the functional derivative considered.

Self-assembled monolayers (SAMs) can be obtained easily and spontaneously, and their preparation is recognized to be compatible with the immobilization of a wide variety of functional molecules in a single step. By this route, electroactive molecules derived from (2,2,6,6-tetramethylpiperidin-1-yl)oxyl (TEMPO),⁴ ferrocene,^{5–9} tetrathiafulvalene,^{10–12} thienylene vinylene,¹³ anthraquinone,¹⁴ perylene diimide,^{15,16} or metal complexes¹⁷ for example, could be immobilized on gold surfaces. Likewise, organosulfide modified biological material such as DNA^{18–21} or proteins^{22,23} has also been used to effectively decorate gold substrates. However, the major drawback inherent in the use of SAMs is their moderate stability,²⁴ which limits the application scope, especially as platforms for sensing.²⁵

Methods giving rise to more robust functional surfaces, such as the reduction of diazonium salts, make it possible to propose alternatives.²⁶ Unfortunately, the formation of thick and uncontrolled disorganized layers linked to the strong reactivity of the radicals produced during diazonium reduction remains an

obstacle to the generalization of the method towards the development of optimized and controlled functional devices.²⁷ In addition, the electrochemical methods generally required to reduce the precursors undeniably add complexity to the implementation, compared to spontaneous methods. Indeed, in addition to a need to adjust various experimental variables (static or dynamic technique, potential, scan rate), the electrografting of fragile entities often requires the implementation of a two-step methodology involving first of all the electro-generation of a reactive film and then the chemical post-functionalization of the platform formed, by the species of interest, under the appropriate conditions.^{28,29}

With the objective of eliminating compromise between control, stability and ease of implementation, the work presented in this article shows that it is possible to spontaneously form monolayers on gold exploiting the diazonium chemistry. Following the model of the SAM precursors, composed of an anchoring head, a spacer and a functional group, we synthesized a molecule comprising the following key elements: an aryldiazonium group (anchoring head), a 12-carbon alkyl chain (spacer) and an electroactive group (ferrocene, Fc) (Figure 1, **DC₁₂Fc**) and investigated its adsorption on gold. The presence of a long alkyl chain is essential since it is expected that it could suppress the electron tunneling through the first grafted organic layer, thus stopping the film growth to produce compact monolayers.³⁰ To validate our approach and justify the hypothesis stated above, a second Fc derivative having a shorter 2-carbon chain was synthesized (Figure 1, **DC₂Fc**) and



immobilized on gold.

Figure 1. DC₁₂Fc and DC₂Fc derivatives.

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Electronic Supplementary Information (ESI) available: Experimental procedures, synthesis of AC₂Fc compound, comparative voltammetric study of bare Fc and DC₁₂Fc, multiple scan rate CV experiments, XPS analyses, AFM analyses of C₂Fc layers. See DOI: 10.1039/x0xx00000x

DC₁₂Fc and DC₂Fc were obtained from the corresponding aniline derivatives (AC₁₂Fc and AC₂Fc, respectively). The aniline derivative AC₁₂Fc was synthesized according to a procedure already described in ref³⁰ and AC₂Fc was synthesized according to a protocol detailed in Supplementary Information.

Gold electrodes obtained by physical vapor deposition (PVD)³¹ were immersed for various times in a solution containing the diazonium salt DC₁₂Fc, formed in-situ from the corresponding aniline (AC₁₂Fc). It is now understood that organic layers obtained by the latter method, conventionally used in many studies, have the same properties as those obtained from isolated diazonium cations.^{32,33} Figure 2 represents the cyclic voltammograms (CVs) recorded on the C₁₂Fc surfaces obtained by immersion of PVD gold electrodes ($S = 0.2 \text{ cm}^2$) in a solution containing 1 mM of AC₁₂Fc and 5 equivalents of *tert*-butyl nitrite (tBuONO) as nitrosating reagent, in dichloromethane.

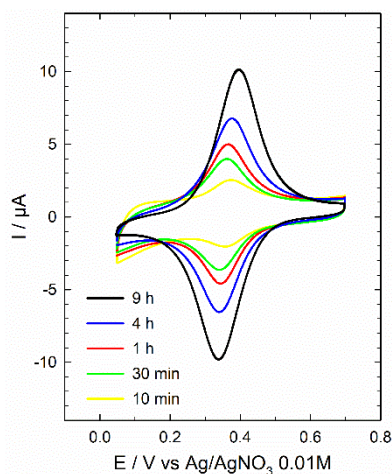


Figure 2. CVs recorded in 0.1 M nBu₄NPF₆ in dichloromethane on layers spontaneously elaborated by immersion during different times in DC₁₂Fc (formed in-situ from 1 mM AC₁₂Fc + 5 mM tBuONO) in dichloromethane. Scan rate: 100 mV.s⁻¹.

All the CVs recorded on Fc-based films obtained after different immersion times show the characteristic reversible one-electron oxidation of ferrocene. The signal intensity increases as the functionalization time increases. Integration of the anodic peaks leads to surface coverages ranging from $7.5 \pm 0.2 \times 10^{-11}$ to $5.5 \pm 0.3 \times 10^{-10} \text{ mol.cm}^{-2}$ for immersion times between 10 min and 9 h. It should be noted that longer functionalization times (overnight) do not lead to higher surface coverages. Interestingly, the threshold value of $5.5 \times 10^{-10} \text{ mol.cm}^{-2}$ matches that previously obtained during a study carried out on Fc-based SAMs on gold,³⁴ which suggests the formation of very thin layers.

To ensure that this value is representative of the entire layer and not linked to incomplete electroactivity of a thick layer, atomic force microscopy (AFM) scratching measurements were performed to determine the thickness of the obtained layers.³⁵ The obtained scratch profile is exploited to extract the layer thickness from the whole scratched area to have an optimal statistic. All experiments were performed under ambient conditions on modified Pyrolyzed Photoresist Films (PPF) characterized by low surface roughness ($< 0.5 \text{ nm}$) in

comparison with gold substrates ($\approx 1 \text{ nm}$). Surface coverages values obtained on PPF substrates were lower than those obtained on PVD gold substrates ($3.9 \times 10^{-10} \text{ mol.cm}^{-2}$ instead of $5.5 \times 10^{-10} \text{ mol.cm}^{-2}$), which is attributed to the lower surface roughness of PPF. Figure 3 shows the topographic image of a scratch on C₁₂Fc layers obtained after a 9 h immersion in the grafting solution (A) and the corresponding depth profile (B).

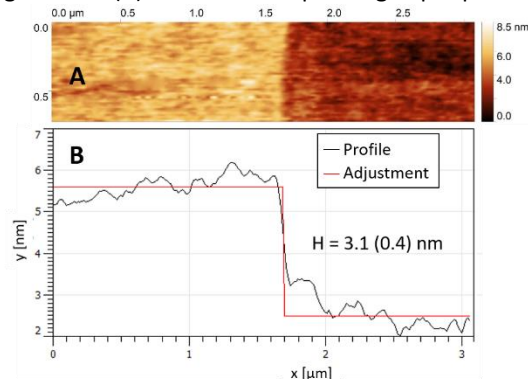


Figure 3. A. AFM image of a C₁₂Fc layer obtained by a 9 h immersion of a PPF substrate, after scratching. B. Depth profile of the scratched area.

An average film thickness of $3.1 \pm 0.4 \text{ nm}$ was estimated from the measurements. Such a value is consistent with the calculated height of a C₁₂Fc group (2.8 nm, calculated by molecular mechanics (MM2) calculations), confirming that the film growth is restricted to an ultrathin layer. Taken together, voltammetric and AFM experiments on C₁₂Fc layers formed from DC₁₂Fc leave no doubt that thin films were obtained.

As reported by Tanaka et al,³⁰ the obtaining of such thin layers can be justified by the following hypothesis: when a first layer of organic C₁₂Fc molecules completely covers the electrode, the tunneling through the formed layer is no longer possible due to the presence of the long alkyl chains. Thus, the reduction of the diazonium cations located at the substrate-solution interface no longer takes place. As no radicals are produced, the film no longer grows and stops at the monolayer stage. To prove this hypothesis, a reducing agent was added to the grafting solution after the formation of a monolayer to confirm that further grafting can occur on the layer via chemical reduction of the diazonium derivative in solution. To this end, bare ferrocene (Fc) was chosen as reducing agent to produce reactive radicals from diazonium cations. Although the DC₁₂Fc layer precursor already contains a Fc entity, the latter cannot act itself as a reducing agent because its cyclopentadiene ring substitution by an amide withdrawing function increases the Fc oxidation potential. Indeed, a difference of 160 mV was measured between the redox potential of bare Fc and that of Fc carried by the DC₁₂Fc molecule, preventing the reduction of the diazonium function (see Figure SI-2).

A monolayer was first formed by a 24 hour-immersion in a solution containing DC₁₂Fc in-situ generated from AC₁₂Fc and tBuONO. This C₁₂Fc-functionalized gold electrode was then immersed in a second solution containing, in addition to AC₁₂Fc and tBuONO, bare Fc as a reducing agent (Figure 4).

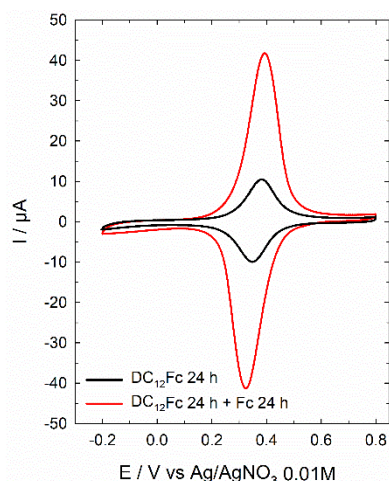


Figure 4. CVs recorded in 0.1 M $n\text{Bu}_4\text{NPF}_6$ in dichloromethane on C_{12}Fc layers spontaneously elaborated by 24 h-immersion in DC_{12}Fc (formed in-situ from 1 mM AC_{12}Fc + 5 mM tBuONO) in dichloromethane (black line) and by subsequent 24h-immersion in a second solution containing DC_{12}Fc (formed in situ from 1 mM AC_{12}Fc + 5 mM tBuONO) and 1 mM bare Fc in dichloromethane (red line). Scan rate: $100 \text{ mV}\cdot\text{s}^{-1}$.

The CV recorded on a C_{12}Fc layer spontaneously obtained from DC_{12}Fc after a 24 h-immersion ($I_{\text{peak}} = 11 \mu\text{A}$) is similar to the one recorded after a 9 h-immersion (Figure 2, $I_{\text{peak}} = 11 \mu\text{A}$). After a subsequent 24 h-immersion in a grafting solution containing bare Fc, the surface coverage had increased 5-fold. The addition of a reducing agent in the presence of the in-situ formed diazonium cation therefore allowed the additional grafting of a large quantity of aryl entities.

This observation proves on the one hand that the reduction of the diazonium function is possible under the experimental conditions used and, on the other hand, that the radicals formed during this reduction are able to react with the monolayer already formed. This experiment thus validates the initial hypothesis, which is based on the inhibition of the diazoniums reduction by the surface when the latter is covered with a layer composed of C_{12}Fc entities. The addition of a reducing agent is necessary for this reduction to occur again.

A second experiment was carried out to highlight the role of the 12-carbon chain in suppressing the tunnel effect through an already grafted C_{12}Fc monolayer: the spontaneous grafting of a Fc-functional diazonium incorporating a shorter alkyl chain. CVs presented in Figure 5 show the comparative study of C_{12}Fc and C_2Fc layers obtained by immersion of the electrode in solutions containing DC_{12}Fc or DC_2Fc , respectively.

Both voltammograms show the characteristic redox process of the ferrocene but a much larger amount of immobilized species is obtained for the C_2Fc layer (threshold values at least three times greater). This observation remains valid for all immersion times tested (Figure 5B). The measured value for a 7 h-immersion (i.e. $18 \times 10^{-10} \text{ mol}\cdot\text{cm}^{-2}$) is typical of a multilayer formation and thus demonstrates the possible grafting of aryl radicals on an already grafted monolayer. Multiple-scan rate cyclic voltammetry experiments were carried out on C_{12}Fc and C_2Fc layers in order to evaluate the evolution of the electroactivity of the immobilized species as a function of the scan rate. For the C_{12}Fc layer, the graph representing $\log(I_{\text{peak}})$

vs $\log(\text{scan rate})$ shows a slope close to 1, in agreement with the formation of a monolayer within which all of the immobilized species remain electroactive, even at high scan rate. On the opposite, data recorded on C_2Fc layers show a lower slope (0.68) owing to the formation of multilayered species (Figure SI-3)

These results show that the reduction of aryldiazonium salts in solution (by electron transfer from the surface through the organic layer) becomes possible when immobilized molecules carry an alkyl chain of only 2 carbons.

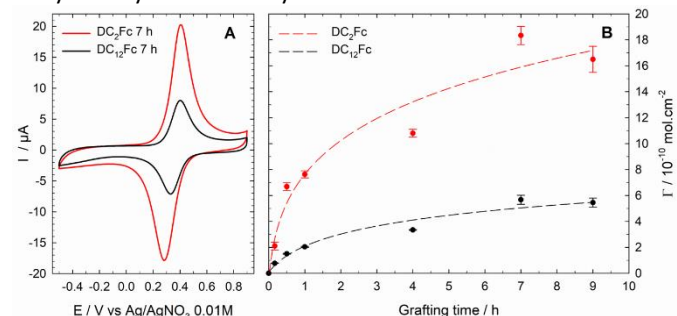


Figure 5. A. CVs recorded in 0.1 M $n\text{Bu}_4\text{NPF}_6$ in dichloromethane on layers spontaneously elaborated from DC_{12}Fc (black line) or from DC_2Fc (red line), by 7 h-immersion in 1 mM aniline derivative + 5 mM tBuONO in dichloromethane. Scan rate: $100 \text{ mV}\cdot\text{s}^{-1}$. B. Evolution of the Fc surface coverage as a function of grafting time, the dotted curves are not trend curves but visual references.

X-ray Photoelectron Spectroscopy (XPS) measurements undertaken on C_{12}Fc and C_2Fc layers confirm the results obtained by cyclic voltammetry. Indeed, the higher intensity of the $\text{Fe}2\text{p}$ signal corresponding to the Fc entities, recorded on C_2Fc layer, highlighted a larger amount of immobilized species, as demonstrated by electrochemical studies (see Figure SI-4A). Moreover, considering the depth of the sample probed (about 10 nm), the signal related to the substrate ($\text{Au}4\text{f}$) is all the more intense as the film is thin. This is the reason why the $\text{Au}4\text{f}$ signal recorded on C_{12}Fc film is higher than the one recorded on C_2Fc film, which is thicker (see Figure SI-4B).

AFM analyses performed on C_2Fc films did not allow determination of the layer thickness since the heterogeneity of the obtained films led to very disparate values (see Figure SI-5). A comparative stability study was then carried out on SAMs elaborated from the thiol derivative $\text{SH-C}_{15}\text{-Fc}$ (structure drawn on Figure 6B and synthesis described in ref³⁴) and layers obtained from DC_{12}Fc diazonium. Both monolayers were synthesized under the same conditions, by a 7 h-immersion of a gold electrode in a solution containing 1 mM of precursor (i.e. DC_{12}Fc formed in-situ or $\text{SH-C}_{15}\text{-Fc}$). After rinsing with dichloromethane, the modified electrodes were studied by

repetitive cyclic voltammetry in the same potential range (0 V - 0.8 V vs Ag/AgNO₃ 0.01 M, Figure 6A).

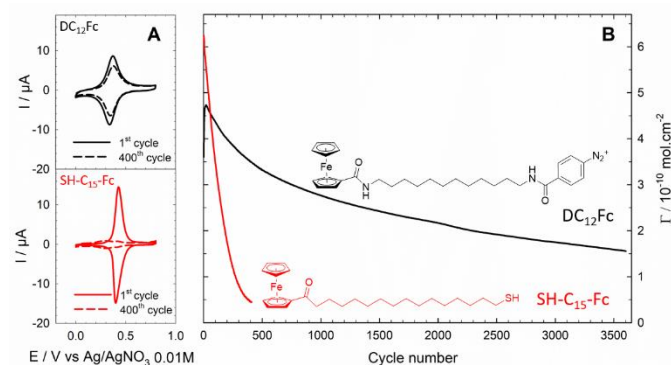


Figure 6. A. Top. CVs (20th cycle: solid line and 400th cycle: dashed line) recorded in 0.1 M nBu₄NPF₆ in dichloromethane on layers spontaneously elaborated by 7 h-immersion in 1 mM DC₁₂Fc formed in-situ (black). Bottom. CVs (20th cycle: solid line and 400th cycle: dashed line) recorded in 0.1 M nBu₄NPF₆ in dichloromethane on layers spontaneously elaborated by 7 h-immersion in 1 mM SH-C₁₅-Fc (red). Scan rate: 100 mV.s⁻¹. B. Evolution of the Fc surface coverage vs cycle numbers.

Both CVs presented in Figure 6A show the characteristic ferrocene signal, located around 0.4 V but significant differences have to be noted. First, the weaker full width at half maximum (FWHM) for the voltammetric peak recorded on the SAM (75 mV versus 120 mV for the C₁₂Fc layer) shows that the structures of the two layers are not strictly identical. Second, the surface coverage, which goes from 5.6 × 10⁻¹⁰ mol.cm⁻² to a null value for the SAM between the 20th and the 400th cycle, shows the total disappearance of the signal after this electrochemical bias. On the contrary, the C₁₂Fc layer unambiguously shows a drastic stability increase since calculated surface coverages corresponding to 74% and 33% of the value recorded for the 20th cycle are found after 400 and 3600 cycles respectively (Figure 6B).

These series of experiments clearly demonstrate the advantage provided by the surface modification from diazonium salts in terms of sustainability. The most likely explanation lies in the lability of the sulfur-gold bond which results in an equilibrium between the immobilized species and the SAM storage medium, thus promoting desorption of the species. The higher robustness of the carbon-gold bond explains the lower signal loss recorded on the layer obtained from DC₁₂Fc and the charge decrease observed during the potential scans can be related to an electrochemical deactivation of the immobilized species or corrosion of gold substrate.

To expand the application range of surfaces modified by organic molecules, it is essential to move towards methods that are both simple to implement and give rise to stable and controlled devices. The methodology proposed here appears to meet these criteria. Indeed, the incorporation of a long alkyl chain between the aryl diazonium and the functional moiety allows growth to be terminated at the monolayer stage while taking advantage of the stability brought by diazonium grafting.

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

There are no conflicts to declare.

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Notes and references

- J. J. Gooding and V. R. Gonçalves, *Curr. Opin. Electrochem.*, 2017, **5**, 203–210.
- T. Wu, C. M. Fitchett, P. A. Brooksby and A. J. Downard, *ACS Appl. Mater. Interfaces*, 2021, **13**, 11545–11570.
- D. Li, Y. Luo, D. Onidas, L. He, M. Jin, F. Gazeau, J. Pinson and C. Mangeney, *Adv. Colloid Interface Sci.*, 2021, **294**, 102479.
- C. Gautier, O. Alévêque, F. Seladji, M. Dias, T. Breton and E. Levillain, *Electrochem. Commun.*, 2010, **12**, 79–82.
- B. Liu, A. J. Bard, M. V. Mirkin and S. E. Creager, *J. Am. Chem. Soc.*, 2004, **126**, 1485–1492.
- P.-Y. Blanchard, S. Boisard, M. Dias, T. Breton, C. Gautier and E. Levillain, *Langmuir*, 2012, **28**, 12067–12070.
- R. A. Wong, Y. Yokota, M. Wakisaka, J. Inukai and Y. Kim, *Nat. Commun.*, 2020, **11**, 4194.
- S. C. Patrick, R. Hein, M. Sharafeldin, X. Li, P. D. Beer and J. J. Davis, *Chem. – Eur. J.*, 2021, **27**, 17700–17706.
- R. M. Mayall, A. J. Marengo, M. Kilgore, V. I. Birss and S. E. Creager, *ChemElectroChem*, 2021, **8**, 1873–1880.
- P.-Y. Blanchard, O. Aleveque, S. Boisard, C. Gautier, A. El-Ghayoury, F. Le Derf, T. Breton and E. Levillain, *Phys. Chem. Chem. Phys.*, 2011, **13**, 2118–2120.
- J. Casado-Montenegro, M. Mas-Torrent, F. Otón, N. Crivillers, J. Veciana and C. Rovira, *Chem. Commun.*, 2013, **49**, 8084.
- C. Adam, L. Faour, V. Bonnin, T. Breton, E. Levillain, M. Sallé, C. Gautier and D. Canevet, *Chem. Commun.*, 2019, **55**, 8426–8429.
- S. Bkhach, O. Alévêque, P. Blanchard, C. Gautier and E. Levillain, *Nanoscale*, 2018, **10**, 1613–1616.
- J. Casado-Montenegro, E. Marchante, N. Crivillers, C. Rovira and M. Mas-Torrent, *ChemPhysChem*, 2016, **17**, 1810–1814.
- S. Bkhach, Y. Le Duc, O. Alévêque, C. Gautier, P. Hudhomme and E. Levillain, *ChemElectroChem*, 2016, **3**, 887–891.
- S. Bkhach, O. Alévêque, Y. Morille, T. Breton, P. Hudhomme, C. Gautier and E. Levillain, *ChemElectroChem*, 2017, **4**, 601–606.
- J. Park, L. Belding, L. Yuan, M. P. S. Mousavi, S. E. Root, H. J. Yoon and G. M. Whitesides, *J. Am. Chem. Soc.*, 2021, **143**, 2156–2163.
- D. A. D. Giusto, W. A. Wlassoff, S. Giesebrecht, J. J. Gooding and G. C. King, *Angew. Chem. Int. Ed.*, 2004, **43**, 2809–2812.
- M. Castronovo, A. Lucesoli, P. Parisse, A. Kurnikova, A. Malhotra, M. Grassi, G. Grassi, B. Scaggiante, L. Casalis and G. Scoles, *Nat. Commun.*, 2011, **2**, 297.
- J. D. Slinker, N. B. Muren, S. E. Renfrew and J. K. Barton, *Nat. Chem.*, 2011, **3**, 228–233.

- 21 D. Kang, J. Yu, F. Xia, J. Huang, H. Zeng, M. Tirrell, J. Israelachvili and K. W. Plaxco, *Langmuir*, 2021, **37**, 13346–13352.
- 22 W. Lee, B.-K. Oh, Y. Min Bae, S.-H. Paek, W. Hong Lee and J.-W. Choi, *Biosens. Bioelectron.*, 2003, **19**, 185–192.
- 23 I. Willner and E. Katz, *Angew. Chem. Int. Ed.*, 2000, **39**, 1180–1218.
- 24 J. C. Love, L. A. Estroff, J. K. Kriebel, R. G. Nuzzo and G. M. Whitesides, *Chem Rev*, 2005, **105**, 1103–1170.
- 25 G. Liu, T. Bocking and J. J. Gooding, *J. Electroanal. Chem.*, 2007, **600**, 335–344.
- 26 A. Mattiuzzi, I. Jabin, C. Mangeney, C. Roux, O. Reinaud, L. Santos, J.-F. Bergamini, P. Hapiot and C. Lagrost, *Nat. Commun.*, 2012, **3**, 1130.
- 27 D. Belanger and J. Pinson, *Chem. Soc. Rev.*, 2011, **40**, 3995–4048.
- 28 N. Marshall, A. Rodriguez and S. Crittenden, *RSC Adv.*, 2018, **8**, 6690–6698.
- 29 C. Gautier, I. López and T. Breton, *Mater. Adv.*, 2021, **2**, 2773–2810.
- 30 M. Tanaka, T. Sawaguchi, Y. Sato, K. Yoshioka and O. Niwa, *Langmuir*, 2011, **27**, 170–178.
- 31 O. Alévêque, P.-Y. Blanchard, T. Breton, M. Dias, C. Gautier and E. Levillain, *Electrochem. Commun.*, 2012, **16**, 6–9.
- 32 S. Baranton and D. Bélanger, *J. Phys. Chem. B*, 2005, **109**, 24401–24410.
- 33 G. Liu, M. Chockalingham, S. M. Khor, A. L. Gui and J. J. Gooding, *Electroanalysis*, 2010, **22**, 918–926.
- 34 P.-Y. Blanchard, G. T. Kenfack, E. Levillain and C. Gautier, *ChemistrySelect*, 2016, **1**, 3171–3174.
- 35 P. A. Brooksby and A. J. Downard, *Langmuir*, 2004, **20**, 5038–5045.